

## Pathology

### MULTIPLE CHOICE:

- The hallmark of acute inflammation is:
  - Increased blood flow
  - Rubor, calor, dolor increased permeability of microcirculation
  - Increased vascular permeability**
  - Vascular stasis
- The most efficient bactericidal compound present in neutrophils is the:
  - Arachidonic acid metabolites
  - H<sub>2</sub>O<sub>2</sub>-MPO-Halide system
  - Lysosomal enzymes**
  - Hydrogen peroxide
- Leukotrienes are derived from arachidonic acid through the help of this enzyme:
  - Phospholipase
  - Lipoxygenase
  - Cyclooxygenase**
  - Lipoxins
- The hallmark of chronic inflammation is:
  - Increased permeability of microcirculation
  - Migration of macrophages at site of injury
  - Tissue destruction**
  - Mononuclear cell infiltration
- An example of tumor suppressor gene:
  - p53**
  - C-myc
  - ras
  - bcr
- A lymph node biopsy was diagnosed as tuberculosis. This type of necrosis is characteristic of:
  - Coagulative
  - Liquefactive
  - Caseous**
  - Gangrenous
- Which of the following substances accumulates in atherosclerotic blood vessels?
  - Fatty acids
  - Cholesterol esters**
  - Triglycerides
  - Phospholipids
- "Heart failure cells" found in the lungs in Chronic Passive Congestion are actually.....
  - Lymphocytes
  - Eosinophils
  - Macrophages**
  - Neutrophils
- The outcome of thrombosis wherein the thrombi tends to GROW due to deposition of additional platelets, fibrin and red blood cells:
  - Resolution
  - Organization
  - Propagation**
  - Recanalization
- Most common cause of death in burn patients who got into shock is:
  - Sepsis
  - CHF
  - Dehydration**
  - Hemorrhage
- This type of hypersensitivity reaction is mediated by IgE:
  - Type I**
  - Type II
  - Type III
  - Type IV
- One of the three major components of acute inflammation include:
  - Alteration in vascular caliber that leads to a decrease in blood flow
  - Structure changes in the microvasculature permitting plasma proteins and leukocytes to leave the circulation**
  - Tissue destruction
  - Healing by connective tissue

# Clinical Pathology With Multiple Choice Questions

**Joacim Rocklöv**



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of 8.001 grams of copper(II) sulfate ( $\text{CuSO}_4$ ) and 2.0153 grams of iron powder ( $\text{Fe}$ ) the goal of this experiment was to determine the product of copper(II) sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8.00 grams of copper(II) sulfate onto a piece of filter paper taking the filter paper's mass into account add these to the beaker 3 measure 50.0 ml of water in the cylinder and add it to the beaker 4

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper(II) sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper(II) sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper(II) sulfate in the iron copper(II) sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper(II) sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper(II) sulfate ions the copper(II) is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalise of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated  $\text{CuSO}_4$  solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

**stoichiometry the reaction of iron with copper ii sulfate** - Jul 14 2023

web stoichiometry the reaction of iron with copper ii sulfate purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction

**stoichiometry lab iron and copper sulfate studylib net** - May 12 2023

web stoichiometry lab iron and copper sulfate sample data data and observations mass of copper ii sulfate pentahydrate 3 27 g of g mass of filter paper mass of filter paper and dry copper mass of dry copper 1 28 g