

## ANALYTICAL CHEMISTRY PRACTICE TEST 1

10 items – 20 minutes

- Protium, the most common isotope of hydrogen has one proton and one electron). What is its mass?  
A.  $1.67 \times 10^{-27}$  g  
B.  $3.35 \times 10^{-27}$  g  
C.  $1.67 \times 10^{-27}$  kg  
D.  $3.35 \times 10^{-27}$  kg
- What is the value of the universal gas constant in cal/mol-K?  
A. 1.986  
B. 3.973  
C. 4.184  
D. 8.314
- Which of the following is NOT equivalent to 100 pF?  
A.  $1 \times 10^{-10}$  F  
B.  $0.1 \times 10^{-9}$  nF  
C.  $1 \times 10^2$  fF  
D. All are equivalent
- How many significant figures are there in 4.200 g?  
A. Infinite  
B. 2  
C. 3  
D. 4
- Arrange the following scenarios according to the scientific method:  
I. Different solvents are used in trials to process raw material X.  
II. Raw material X yielded more products than the original material.  
III. A previous study suggested the usage of a different raw material X.  
A. III, I, II  
B. I, II, III  
C. III, II, I  
D. I, III, II

For numbers 6-7: Consider the experimental results of the effect of solvent in the amount of phytochemical extracted.

- How much solvent will be required to extract 1.2 g of phytochemical?  
A. 38 mL  
B. 43 mL  
C. 51 mL  
D. 55 mL
- Which of the following can be done in order to extract 2.4 g of phytochemical?  
A. Use 80 mL of solvent in one run.  
B. Use 40 mL of solvent in two runs.  
C. Both A and B.  
D. None of the above.
- Which of the following is an extensive property?  
A. melting point  
B. density  
C. electric potential  
D. none of these
- A colorless liquid and a yellowish liquid were mixed in a test tube. Which of the following is incorrect?  
A. There will be no reaction if the test tube will get hot.  
B. There will be a reaction if a precipitate will be formed.  
C. There will be no reaction if the liquids will be on top of each other.  
D. There will be a reaction if the yellow color will drastically disappear.
- Which of the following is TRUE?  
A. Compounds have only one type of atom.  
B. Elements can be chemically decomposed.  
C. Solutions are considered as pure substances.  
D. Heterogeneous mixtures have phase boundaries.



# Analytical Chemistry Practice

**Léo Pomar**



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