

If r_1 and r_2 be the rates of diffusion of two gases A and B, d_1 and d_2 be their densities. Then, According to Graham's law of diffusion

$$r_1 \propto \frac{1}{\sqrt{d_1}} \quad \& \quad r_2 \propto \frac{1}{\sqrt{d_2}}$$

$$\therefore r_1 = K \frac{1}{\sqrt{d_1}} \quad \& \quad r_2 = K \frac{1}{\sqrt{d_2}} \quad \text{where 'K' is proportionality}$$

constant

$$\text{Then,} \quad \frac{r_1}{r_2} = \frac{\sqrt{d_2}}{\sqrt{d_1}}$$

$$\frac{r_1}{r_2} = \frac{\sqrt{2 \times d_2}}{\sqrt{2 \times d_1}}$$

$$\frac{r_1}{r_2} = \frac{\sqrt{M_2}}{\sqrt{M_1}} \quad (\text{Mol. Mass, } M = 2 \times V.D.)$$

Where M_1 & M_2 is molecular mass of A & B respectively.

Thus, Graham's law may also be stated as, "Under similar condition of temperature and pressure, the rate of diffusion of gases are inversely proportional to square root of their molecular masses."

$$\text{Rate of diffusion} = \frac{\text{Volume of gas diffused}}{\text{Time taken to diffuse}}$$

$$r_1 = \frac{V}{t_1}$$

$$r_2 = \frac{V}{t_2}$$

$$\therefore \frac{r_1}{r_2} = \frac{V}{t_1} \times \frac{t_2}{V} = \frac{t_2}{t_1}$$

$$\therefore \frac{t_2}{t_1} = \frac{\sqrt{d_2}}{\sqrt{d_1}} = \frac{\sqrt{M_2}}{\sqrt{M_1}}$$

The laws of diffusion are applicable to Effusion also. Effusion is a process in which a gas under pressure escapes out of a fine hole e.g. Puncture in tube of tyres is a process of effusion. Effusion is always followed by diffusion.

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Robert N. Schock



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The Interface Structure and Electrochemical Processes at the Boundary Between Two Immiscible Liquids Vladimir E. Kazarinov, 2012-12-06 Studies on the electrochemical processes at the interface between two immiscible liquids began a long time ago they date back to the end of the last century Such celebrated scientists as Nemst and Haber and also young A N Frumkin were among those who originated this science Later A N Frumkin went a long way in furthering the studies at the Institute of Electrochemistry The theory of the appearance of potential in a system of two immiscible electrolytes was developed and experimentally verified before the beginning of the thirties In later years the studies in this area considerably lagged behind those conducted at metal electrodes which were widely used in different industries In the past 15 years however the situation has radically changed and we have witnessed a drastic increase in the number of publications on the electrochemistry of immiscible electrolytes We are glad to note that the investigations show not only a quantitative but also a qualitative change The theoretical works on the oil water interface test not only the thermodynamic aspects of the interface but also recreate the molecular picture of the process Along with the now conventional oilwater system electrochemical studies are made on various membranes including the finest bilayer lipid membranes and also on microemulsion systems A prominent place in the investigation of the oil water interface is occupied by photoprocesses that come into play at the interface between two ionic conductors

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