

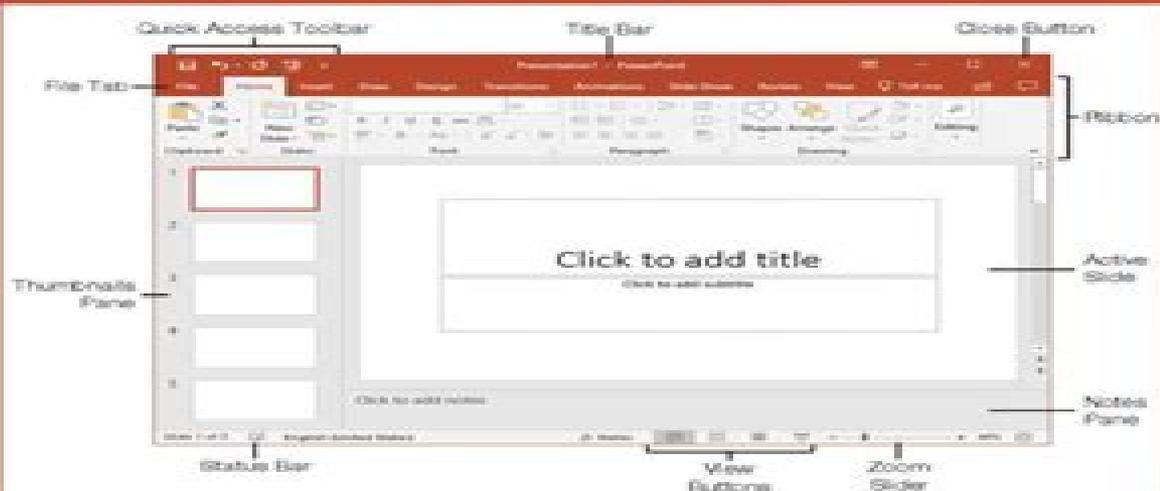


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The PowerPoint 2016 Program Screen



Keyboard Shortcuts

General

Open a presentation.....	Ctrl + O
Create a new presentation	Ctrl + N
Save a presentation.....	Ctrl + S
Print a presentation.....	Ctrl + P
Insert a new slide.....	Ctrl + M
Toggle the Notes pane.....	Ctrl + Shift + H
Close a presentation	Ctrl + W
Exit PowerPoint.....	Ctrl + Q
Help.....	F1

Editing

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Paste.....	Ctrl + V
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Redo.....	Ctrl + Y
Group.....	Ctrl + G
Find.....	Ctrl + F
Replace.....	Ctrl + H
Select All.....	Ctrl + A

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Bold.....	Ctrl + B
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Underline.....	Ctrl + U
Align Left.....	Ctrl + L
Align Right.....	Ctrl + R
Center.....	Ctrl + E
Justify.....	Ctrl + J

Slide Show Delivery

Begin slide show.....	F5
Resume slide show.....	Shift + F5
End slide show.....	Esc
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Go to previous slide.....	Page Up
Go to first slide.....	Home
Go to last slide.....	End
Jump to slide.....	Slide # + Enter
Toggle screen black.....	B
Toggle screen white.....	W
Pause show.....	S
Change arrow to pen.....	Ctrl + P
Change pen to arrow.....	Ctrl + A

Getting Started



The **File** tab opens Backstage view, which contains commands for working with your files like **Open**, **Save**, **New**, **Print**, **Share**, and **Close**.

Create a New Presentation: Click the **File** tab, select **New**, and click **Blank Presentation**. Or, press **Ctrl + N**.

Open a Presentation: Click the **File** tab and select **Open**, or click **Ctrl + O**. Select a location with a file you want, then select a file and click **Open**.

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Undo: Click the **Undo** (↶) button on the Quick Access Toolbar.

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- Normal View**
- Outline View**
- Slide Sorter View**
- Reading View**

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8.00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper's mass into account add these to the beaker 3 measure 50.0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in

our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalization of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7.0002g of copper ii sulfate and 2.0101g of iron metal which produced 2.3037g of copper product with a percent yield of 100.72% introduction

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