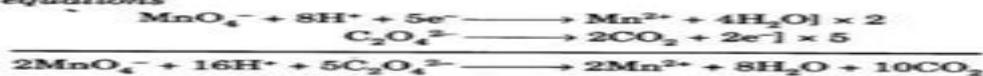


**Objectives:**

A. Standardization of KMnO_4 solution by using a primary standard 0.1M Oxalic acid solution:

Ionic equations

**LAB REPORT:**

Calculations:

Trial	Start Volume	End Volume	Volume taken
Trail 1			
Trail 2			
Trail 3			
Average			

At the end point:

$\frac{\text{The number of moles of KMnO}_4}{\text{The number of moles of H}_2\text{C}_2\text{O}_4} = \text{mole ratio}$

$$\frac{MV (\text{KMnO}_4)}{MV (\text{H}_2\text{C}_2\text{O}_4)} = \frac{2}{5}$$

$$\frac{M \times \text{average volume} (\text{KMnO}_4)}{0.05 \times 10 (\text{H}_2\text{C}_2\text{O}_4)} = \frac{2}{5}$$

$$M_{\text{KMnO}_4} = 2(0.05 \times 10) / (5 \times \text{average volume}) = \quad \text{mol/L}$$

Notes:

1. Permanganate is a self-indicator
2. Permanganate titrations must be carried out in an acidic medium, or else, a brown precipitate of MnO_2 will be formed which prevents clear endpoint

Now, after determination of the permanganate molarity, we can use it to determine unknown Fe (II) solution as follows:

Objectives:

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L Manion



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