

# Some Basic Principles of Chemistry

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1. Suppose the elements  $X$  and  $Y$  combine to form two compounds  $XY_2$  and  $X_3Y_2$ . When 0.1 mole of  $XY_2$  weighs 10 g and 0.05 mole of  $X_3Y_2$  weighs 9 g, the atomic weights of  $X$  and  $Y$  are [NEET 2016 Phase II]  
(a) 40, 30 (b) 60, 40 (c) 20, 30 (d) 30, 20
2. If Avogadro number  $N_A$  is changed from  $6.022 \times 10^{23} \text{ mol}^{-1}$  to  $6.022 \times 10^{20} \text{ mol}^{-1}$  this would change [CBSE AIPMT 2015]  
(a) the definition of mass in units of grams  
(b) the mass of one mole of carbon  
(c) the ratio of chemical species to each other in a balanced equation  
(d) the ratio of elements to each other in a compound
3. 20.0 g of a magnesium carbonate sample decomposes on heating to give carbon dioxide and 8.0 g magnesium oxide. What will be the percentage purity of magnesium carbonate in the sample? (Atomic weight of Mg = 24) [CBSE AIPMT 2015]  
(a) 75 (b) 96 (c) 60 (d) 84
4. What is the mass of precipitate formed when 50 mL of 16.9% solution of  $\text{AgNO}_3$  is mixed with 50 mL of 5.8%  $\text{NaCl}$  solution? (Ag = 107.8, N = 14, O = 16, Na = 23, Cl = 35.5) [CBSE AIPMT 2015]  
(a) 28 g (b) 3.5 g (c) 7 g (d) 14 g
5. The number of water molecules is maximum in [CBSE AIPMT 2015]  
(a) 18 molecules water (b) 1.8 g of water  
(c) 18 g of water (d) 18 moles of water
6. When 22.4 L of  $\text{H}_2(\text{g})$  is mixed with 11.2 L of  $\text{Cl}_2(\text{g})$ , each at STP, the moles of  $\text{HCl}(\text{g})$  formed is equal to [CBSE AIPMT 2014]  
(a) 1 mole of  $\text{HCl}(\text{g})$  (b) 2 moles of  $\text{HCl}(\text{g})$   
(c) 0.5 mole of  $\text{HCl}(\text{g})$  (d) 1.5 moles of  $\text{HCl}(\text{g})$
7. 1.0 g of magnesium is burnt with 0.56 g of oxygen in a closed vessel. Which reactant is left in excess and how much? [CBSE AIPMT 2014]  
(At. weight of Mg = 24, O = 16)  
(a) Mg, 0.16 g (b)  $\text{O}_2$ , 0.16 g  
(c) Mg, 0.44 g (d)  $\text{O}_2$ , 0.28 g
8. How many grams of concentrated nitric acid solution should be used to prepare 250 mL of 2.0 M  $\text{HNO}_3$ ? The concentrated acid is 70%  $\text{HNO}_3$ . [NEET 2013]  
(a) 45.0 g conc.  $\text{HNO}_3$  (b) 90.0 g conc.  $\text{HNO}_3$   
(c) 70.0 g conc.  $\text{HNO}_3$  (d) 54.0 g conc.  $\text{HNO}_3$
9.  $6.02 \times 10^{20}$  molecules of urea are present in 100 mL of its solution. The concentration of solution is [NEET 2013]  
(a) 0.02 M (b) 0.01 M  
(c) 0.001 M (d) 0.1 M

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**A Gutmann**



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