

Chemistry

Basic Concepts of chemistry

- Density= mass/volume
- Celsius to Fahrenheit conversion: $^{\circ}\text{F} = 9/5(^{\circ}\text{C}) + 32$
- Celsius to Kelvin conversion: $\text{K} = ^{\circ}\text{C} + 273.15$
- 1 Atomic mass unit(amu)= 1.66056×10^{-24}

-Molarity (M) =
$$\frac{n \text{ (moles of solute)}}{V \text{ (volume of solution)}}$$

-Dilution formula:

$$M_1V_1 = M_2V_2$$

M_1 = initial molarity ("stock solution")

V_1 = initial volume (liters)

M_2 = final (desired) molarity

V_2 = final volume (liters)



-Atomic mass of an element =

$$\frac{\text{mass of one atom of the element}}{\text{mass of } \frac{1}{12} \text{th part of mass of } \text{C}^{12} \text{ atom}}$$
$$\frac{\text{mass of one atom of the element}}{\text{mass of } \frac{1}{16} \text{th part of mass of } \text{O}^{16} \text{ atom}}$$
$$\frac{\text{mass of one atom of the element}}{\text{mass of one atom of Hydrogen atom}}$$

-Relative Atomic mass =

$$\frac{\text{Average mass of atoms of an element}}{\text{Mass of one atom of carbon-12}} \times 12$$

-Number of molecules in n moles of substance = $n \times N_A$

-Mass percentage of an element in a compound =

$$\frac{\text{mass of that element in the compound}}{\text{molar mass of the compound}} \times 100$$

-Mass percent = $\frac{\text{Mass of the solute}}{\text{Mass of the solution}} \times 100$

-Avogadro's No = 6.022×10^{23}

-Molecular mass = 2 * vapour density

-Mole fraction of solute = $\frac{\text{moles of solute}}{\text{total moles in solutions}}$

-Molality(m) = $\frac{\text{mole of solute} \times 1000}{\text{volume of solvent in kg}}$

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