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web a write a balanced equation for the reaction Al + O₂ → Al₂O₃ b determine the theoretical yield of Al₂O₃ 160.0 g Al × 1 mol Al / 27.0 g Al × 2 mol Al₂O₃ / 3 mol Al × 101.964 g Al₂O₃ / 3 mol Al₂O₃ = 269.82 g Al₂O₃ c determine the percent yield yield actual / yield theoretical × 100 = 260.0 g / 269.82 g × 100 = 96.73%

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web 7Na₂C₂O₄ is the limiting reactant percent yield = 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H₂ are left over 13

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